From: youngg@adelphia.net
Sent: Wednesday, April 05, 2006 2:39 PM
To: Benham, Katherine; zea@ccof.org%inter2; milla@handleycellars.com%inter2
Subject: Comments on hydrated lime

Attachments: ATTACHMENT.TXT
I have the following comments on hydrated lime:

First, it is not produced by the "burning of various forms of limestone". Limestone does not "burn" (Combustion is an oxidation process in which carbon or other materials are oxidized). Limestone CaCO3 is heated, which drives off the C02, leaving CaO. Water is then added to hydrate the lime. This could be considered a chemical change, but then anything that ionizes in water could also be considered as such.

I have recommended the use of CaOH for 25 years with Bordeaux/copper sprays, as a calcium foliar nutrient, and as a stand-alone miticide/fungicide. Although it is high pH and caustic to skin and eyes, it is a relatively safe material to work with.

When CaOH is mixed with water, soil, or air, a certain percentage, and eventually all, would revert to CaCO3 by absorbing CO2 (or its dissolved form, HCO3-

There are finely ground limestones available that might work as substitutes; however these are not registered as pesticide adjuvants at this time.

Hydrated lime could be considered a processed rock powder or a chemically modified mineral and either would be accurate.

I recommend keeping hydrated lime as a tool in organic production.

** For aquatic plant products, there need to be more alternatives allowed for stabilization of kelp. This is a valuable product with many uses in agriculture: root dip, foliar spray, adjuvant for nutrient uptake.

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